

Experimental

Measurements. Electronic absorption spectra were recorded with an Analytikjena Specord 200 UV-vis spectrophotometer, infrared spectra with a Genesis II FT-IR spectrophotometer, NMR spectra with a Varian Mercury 300 FT NMR spectrometer, and conductance measurements with a Metrohm Herisau Conductometer E518. GC-mass spectra were measured with a Shimadzu GCMSD-QP5050 spectrometer. FAB mass spectra were performed at the Korea Basic Science Institute, Daegu, Korea. Elemental analyses were performed at the Research Center for Instrumental Analysis, Daegu University, Gyeongsan, Korea. Magnetic moments were calculated from magnetic susceptibility data obtained at 293 K using a Johnson Matthey MK-1 magnetic susceptibility balance.

Safety Note. Perchlorate salts of metal complexes with organic ligands are often explosive and should be handled with great caution.

Preparation of 17-Cyanomethyl-3,14-dimethyl-2,6,13,17-tetraazatetracyclo[16.4.1^{2.6}.0^{1.18}.0^{7.12}]tricosane (L¹¹) and [NiL²](ClO₄)₂·H₂O (L² = 2-cyanomethyl-5,16-dimethyl-2,6,13,17-tetraazatricyclo[16.4.0.0^{7.12}]docosane). The macropolycycle L¹⁰ was prepared by the reaction of L¹ with formaldehyde.²⁴ A chloroform solution (20 mL) of L¹⁰ (2.0 g, 5.7 mmol) and bromoacetonitrile (0.5 mL, 6.9 mmol) was stirred for 24 h at room temperature and then refluxed for 30 min. The resulting solution was evaporated on a rotary evaporator to dryness. After dissolving the residue in acetonitrile (20 mL), 1.0 M NaOH aqueous solution (20 mL) was added to the solution. The mixture was evaporated at room temperature to precipitate a white solid. The product was collected by filtration and washed with methanol. It was recrystallized by evaporating a chloroform-methanol (1:1) solution (30 mL) of the product. Yield: 50-60%. GC-mass (*m/z*): 386 for L¹¹. Anal. Found: C, 68.97; H, 11.25; N, 18.93. Calc. for C₂₃H₄₁N₅ (L¹¹): C, 71.27; H, 10.66; N, 18.07%. IR (cm⁻¹): 3320 (ν_{N-H}) and 2230 (ν_{C≡N}). Although mass spectrum of the product shows a main peak corresponding to L¹¹, it also contains a small amount of uncharacterized by-product(s). Unfortunately, the isolation of L¹¹ as its pure form was very difficult, and the crude product was used for the preparation of [NiL²](ClO₄)₂·H₂O and [CuL²](ClO₄)₂·H₂O.

An acetonitrile suspension (25 mL) of L¹¹ and Ni(OAc)₂·4H₂O (2.5 g, 10 mmol) was refluxed for 1 h and then cooled to room temperature. After the addition of concentrated HClO₄ (2.0 mL) and water (20 mL), the resulting solution in an open beaker was slowly evaporated at room temperature to precipitate an orange solid. The product was collected by filtration and then dissolved in 1.0 × 10⁻³ M HClO₄ acetonitrile solution (*ca.* 10 mL). The solution was filtered to remove any solid. After the addition of water (10 mL) to the filtrate, the solution was evaporated at room temperature. The orange solid formed was filtered, washed with methanol, and dried in air. Yield: ~40% based on L¹⁰. Anal. Found: C, 40.43; H, 7.44; N, 11.05. Calc. for C₂₂H₄₃N₅Cl₂NiO₉ ([NiL²](ClO₄)₂·H₂O). C, 40.58; H, 6.66; N, 10.75%. FAB

Mass (*m/z*): 531.9 for [NiL² + ClO₄]⁺; 431.9 for [NiL² - H]⁺. IR (cm⁻¹): 3515 (ν_{O-H}, H₂O), 3214 (ν_{N-H}), 3165 (ν_{N-H}), 3087 (ν_{N-H}), 2250 (ν_{C≡N}), and 1620 {δ(H₂O)}. ¹³C-NMR (CD₃NO₂): 16.2, 16.5, 25.2, 25.3, 25.4, 25.7, 26.1, 30.2, 30.7, 30.9, 31.5, 33.1, 39.1, 41.4, 47.3, 48.1, 51.3, 59.2, 66.9, 76.8 (-CH₂CN), 114.5 (-CH₂CN) ppm.

Preparation of [CuL²](ClO₄)₂·H₂O. This complex was prepared by a method similar to that for [NiL²](ClO₄)₂·H₂O except that Cu(OAc)₂·H₂O (2.0 g, 10 mmol) was reacted instead of Ni(OAc)₂·4H₂O. The red-purple solid was recrystallized from 1.0 × 10⁻³ M HClO₄ acetonitrile-water (1:2) solution. Yield: ~40% based on L¹⁰. Anal. Found: C, 40.78; H, 6.88; N, 10.40. Calc. for C₂₂H₄₃N₅Cl₂CuO₉: C, 40.28; H, 6.61; N, 10.67%. FAB Mass (*m/z*): 536.9 for [CuL² + ClO₄]⁺; 438.0 for [CuL² - H]⁺. IR (cm⁻¹): 3515 (ν_{O-H}, H₂O), 3210 (ν_{N-H}), 3185 (ν_{N-H}), 3124 (ν_{N-H}), 2250 (ν_{C≡N}), and 1620 [δ(H₂O)].

Preparation of [Ni(HL³)(H₂O)](ClO₄)₂. The pH of a water-acetonitrile (1:2) solution (20 mL) of [NiL²](ClO₄)₂·H₂O (0.5 g) was adjusted to ≥ 8 through the addition of 0.1 M NaOH solution. The solution was stirred at room temperature for 5 min. The resulting pale purple solution was filtered to remove any solid, and then concentrated HClO₄ (1.0 mL) was added to the filtrate. The mixture was evaporated at room temperature to produce a pale purple solid. The product was collected by filtration, washed with cold methanol, and dried in air. It was recrystallized by the addition of HClO₄ to hot a water-acetonitrile (1:1) solution of the crude product. Yield: ~70%. Anal. Found: C, 39.47; H, 6.51; N, 10.56. Calc. for C₂₂H₄₅N₅Cl₂NiO₁₀: C, 39.48; H, 6.78; N, 10.46%. FAB Mass (*m/z*): 550.2 for [Ni(HL³) + ClO₄]⁺; 450.2 for [Ni(HL³) - H]⁺. IR (cm⁻¹): 3500 (ν_{O-H}, H₂O), 3397 (ν_{N-H}), 3387 (ν_{N-H}), 3340 (ν_{N-H}), 3290 (ν_{N-H}), 3257 (ν_{N-H}), 3238 (ν_{N-H}), 3190 (ν_{N-H}), 1670 (ν_{C=O}), 1620 [δ(H₂O)], and 1600 [δ(NH₂)]. μ_{eff} = 2.80 μ_B.

This complex was also prepared by stirring a warm (~50 °C) basic (pH ≥ 10) water-acetonitrile (1:1) solution (20 mL) of [NiL⁴(H₂O)](ClO₄)₂ (0.5 g) for 10 min. the addition of concentrated of HClO₄ (1.0 mL) to the resulting solution produced the purple solid.

Preparation of [NiL³(H₂O)]ClO₄. The pH of a water-acetonitrile (1:1) solution (20 mL) of [Ni(HL³)(H₂O)](ClO₄)₂ (0.5 g) was adjusted to ≥ 12 through the addition of 1.0 M NaOH solution. The solution was evaporated at room temperature to produce a dark purple solid. The product was collected by filtration, washed with cold water, and dried in air. Yield: ~80%. Found: C, 45.69; H, 7.34; N, 11.92. Calc. for C₂₂H₄₄N₅ClNiO₆: C, 46.46; H, 7.80; N, 12.31%. FAB Mass (*m/z*): 549.9 for [NiL³ + ClO₄]⁺; 450.0 for [NiL³ - H]⁺. IR (cm⁻¹): 3480 (ν_{O-H}, H₂O), 3310 (ν_{N-H}), 3300 (ν_{N-H}), 3252 (ν_{N-H}), 3200 (ν_{N-H}), 1620 [δ(H₂O)], and 1570 {ν_(CONH)}. μ_{eff} = 2.78 μ_B.

Preparation of [Cu(HL³)](ClO₄)₂. This complex was prepared by a method similar to that for [Ni(HL³)(H₂O)](ClO₄)₂ except that [CuL²](ClO₄)₂ (0.5 g) was reacted instead of [NiL²](ClO₄)₂·H₂O. The purple solid was recrystallized from hot acetonitrile-water (1:2) solution. Yield: ~70%. Found: C,

40.90; H, 6.34; N, 10.79. Calc. for $C_{22}H_{43}N_5Cl_2NiO_9$: C, 40.28; H, 6.61; N, 10.67%. FAB Mass (m/z): 554.9 for $[Cu(HL^3) + ClO_4]^+$; 456.0 for $[Cu(HL^3) - H]^+$. IR (cm^{-1}): 3351 (ν_{N-H}), 3320 (ν_{N-H}), 3243 (ν_{N-H}), 3202 (ν_{N-H}), 1680 ($\nu_{C=O}$), and 1600 [$\delta(NH_2)$].

This complex was also prepared by stirring a warm ($\sim 50^\circ C$) basic ($pH \geq 10$) water-acetonitrile (1:1) solution (20 mL) of $[CuL^4](ClO_4)_2$ (0.5 g) for 10 min. The addition of concentrated $HClO_4$ (1.0 mL) to the resulting solution produced the purple solid.

Preparation of $[NiL^4(H_2O)](ClO_4)_2$. To a methanol-acetonitrile (1:3) solution (20 mL) of $[NiL^2](ClO_4)_2 \cdot H_2O$ (0.5 g) was added triethylamine (0.1 mL). The resulting solution was refluxed for 5 min. During which time, the orange solution turned to pale purple. After the addition of concentrated $HClO_4$ (1.0 mL) and water (10 mL) at room temperature, the mixture was evaporated to precipitate a purple solid. The product was collected by filtration, washed with cold methanol, and dried in air. It was recrystallized from hot water-acetonitrile (1:1) mixture. Yield: $\sim 70\%$. Anal. Found: C, 40.90; H, 7.45; N, 10.09. Calc. for $C_{23}H_{47}N_5Cl_2NiO_{10}$: C, 40.43; H, 6.93; N, 10.25%. FAB Mass (m/z): 564.2 for $[NiL^4 + ClO_4]^+$; 464.3 for $[NiL^4 - H]^+$. IR (cm^{-1}): 3500 (ν_{O-H} , H_2O), 3290 (ν_{N-H}), 3250 (ν_{N-H}), 3232 (ν_{N-H}), 1666 ($\nu_{C=N}$), and 1620 [$\delta(H_2O)$]. $\mu_{eff} = 2.76 \mu_B$.

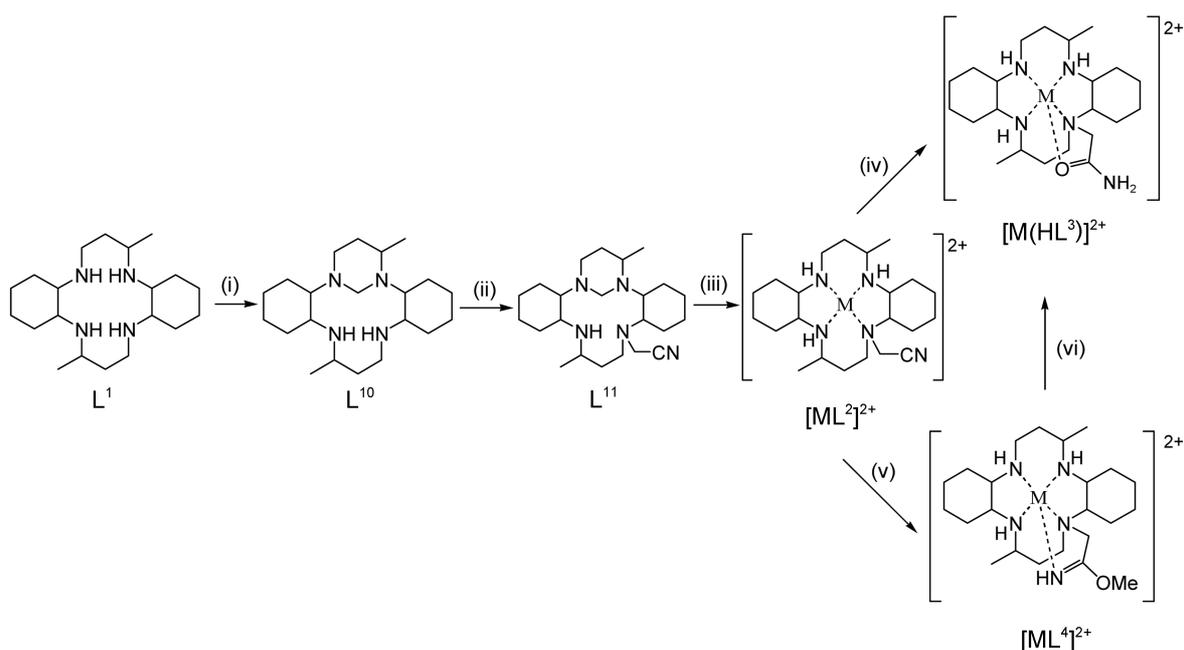
Preparation of $[CuL^4](ClO_4)_2$. This complex was prepared in a methanol-acetonitrile (1:3) solution (20 mL) by a method similar to that for $[NiL^4(H_2O)](ClO_4)_2$ except that $[CuL^2](ClO_4)_2$ (0.5 g) was reacted instead of $[NiL^2](ClO_4)_2 \cdot H_2O$. The purple solid was recrystallized from hot acetonitrile-water (1:2) solution. Yield: $\sim 70\%$. Anal. Found: C, 40.25; H, 7.01; N, 9.93. Calc. for $C_{23}H_{47}N_5Cl_2CuO_{10}$: C, 40.15; H, 6.88; N, 10.18%. FAB Mass (m/z): 569.1 for $[CuL^4 + ClO_4]^+$; 470.2 for $[CuL^4 - H]^+$. IR (cm^{-1}): 3500 (ν_{O-H} , H_2O),

3300 (ν_{N-H}), 3200 (ν_{N-H} , *br*), 1665 ($\nu_{C=N}$), and 1620 [$\delta(H_2O)$].

Results and Discussion

Synthesis. The mono-*N*-hydroxyethylated macrocycle HL^9 has been prepared by the direct reaction of L^1 with $BrCH_2CH_2OH$.¹⁵ However, unexpectedly, our initial attempts to prepare L^2 from the reaction of L^1 with $BrCH_2CN$ in a 1:1 molar ratio were unsuccessful; the solid isolated from the reaction solution was found to be a mixture of L^5 bearing two *N*- CH_2CN pendant arms and the reactant L^1 . Therefore, the synthesis of L^2 begins with the preparation of L^{10} where two amino groups are protected. The reaction of L^{10} with $BrCH_2CN$ produces the macropolycycle L^{11} bearing only one *N*- CH_2CN pendant arm. The macrocycle L^2 could be prepared as its nickel(II) or copper(II) complex, $[ML^2]^{2+}$ ($M = Ni(II)$ or $Cu(II)$), by the reaction of L^{11} with the metal ion in acetonitrile. The nickel(II) and copper(II) complexes of HL^3 and L^4 bearing one *N*- CH_2CONH_2 or *N*- $CH_2C(=NH)OMe$ pendant arm were prepared by the reaction of $[ML^2]^{2+}$ with water or methanol. The reactivity of $[ML^2]^{2+}$ in water or methanol is not quite different from that reported for $[ML^5]^{2+}$.^{17,23} The preparation of nickel(II) and copper(II) complexes of HL^3 can also be achieved by the hydrolysis of the corresponding complexes of L^4 in basic aqueous solutions. The synthetic procedures for the complexes of $L^2 - L^4$ are shown in Scheme 1.

In basic aqueous solutions, $[Ni(HL^3)(H_2O)](ClO_4)_2$ readily undergoes deprotonation to yield $[NiL^3(H_2O)]ClO_4$ ($L^3 =$ a deprotonated form of HL^3). The isolation of $[NiL^3(H_2O)]ClO_4$ as a solid in the present work is quite interesting, because such deprotonation is not observed for other related macrocyclic complexes, such as $[Cu(HL^3)](ClO_4)_2$ and $[NiL^6](ClO_4)_2$. In the case of $[Cu(HL^3)](ClO_4)_2$ or $[NiL^6](ClO_4)_2$, any



Scheme 1. Reagents and Conditions: (i) $HCHO/MeOH$; (ii) $BrCH_2CN/CHCl_3$; (iii) M^{2+}/CH_3CN ; (iv) H_2O ($pH \geq 8$); (v) $MeOH/Et_3N$; (vi) H_2O ($pH \geq 8$). The coordinated solvent molecule is omitted for clarity.

Table 1. Electronic Absorption Spectra and Molar Conductance Data^a

Complex	λ_{\max} , nm (ϵ , M ⁻¹ cm ⁻¹)			Λ_M , (Ω^{-1} cm ² mol ⁻¹)		
[NiL ¹](ClO ₄) ₂ ^b	465(66) ^c					
[NiL ²](ClO ₄) ₂	478(100)	478(82) ^c	477(100) ^d	477 ^e	125	250 ^c 45 ^d
[Ni(HL ³)(H ₂ O)](ClO ₄) ₂		523(7.7)	695(2.6)		110	240 ^c 50 ^d
	335(12) ^c	506(7.4) ^c	724(2.1) ^c			
	345(14) ^d	532(7.5) ^d	700(2.7) ^d			
[NiL ³ (H ₂ O)](ClO ₄) ₂		527(11.6)	740(3.1)		75	130 ^c 24 ^d
	354(42) ^c	556(25) ^c	750(6.7) ^c			
	360(73) ^d	568(40) ^d	750(8.0) ^d			
[NiL ⁴ (H ₂ O)](ClO ₄) ₂		530(20)	710(4.5)		130	233 ^c 55 ^d
	333(13) ^c	515(9.1) ^c	740(3.6) ^c			
	345(16) ^d	532(18) ^d	710(4.5) ^d			
[NiL ⁵](ClO ₄) ₂ ^f	492(77) ^d					
[CuL ¹](ClO ₄) ₂ ^b	487(113) ^c					
[CuL ²](ClO ₄) ₂ ·H ₂ O	497(137)	527(170) ^c	532(165) ^d	495 ^e	120	220 ^c 55 ^d
[Cu(HL ³)](ClO ₄) ₂	540(125)	540(126) ^c	537(130) ^d	545 ^e	130	250 ^c 60 ^d
[CuL ⁴](ClO ₄) ₂	553(120)	552(122) ^c	570(195) ^d	553 ^e	134	235 ^c 56 ^d
[CuL ⁵](ClO ₄) ₂ ^f	535(230) ^c	500 ^e				

^aMeasured in nitromethane at room temperature unless otherwise specified. ^bRef. 24. ^cMeasured in acetonitrile. ^dMeasured in DMSO. ^eMeasured in Nujol mull. ^fRef. 17.

deprotonated form could not be prepared under similar experimental conditions. It is obvious that the acidity of the *N*-CH₂CONH₂ group in [Ni(HL³)(H₂O)](ClO₄)₂ is stronger than that attached to [Cu(HL³)](ClO₄)₂ or [NiL⁶](ClO₄)₂.

Spectra and Properties of [NiL²](ClO₄)₂·H₂O and [CuL²](ClO₄)₂·H₂O. The FAB mass spectrum of [NiL²](ClO₄)₂·H₂O shows two groups of peaks at *m/z* 531.9 {[NiL² + ClO₄]⁺} and 431.9 {[NiL² - H]⁺}. The spectrum of [CuL²](ClO₄)₂·H₂O also shows two groups of peaks corresponding to [CuL² + ClO₄]⁺ and {[CuL² - H]⁺ fragments at *m/z* 536.9 and 438.0, respectively. In the infrared spectra of the complexes, peaks corresponding to $\nu_{\text{N-H}}$ of the coordinated amino groups are observed at 3120-3220 cm⁻¹. The spectra also show $\nu_{\text{C=N}}$ of the pendant arm at *ca.* 2250 cm⁻¹. The nickel(II) complex was found to be a diamagnetic substance. ¹³C-NMR spectrum (*see* Experimental Section) of the nickel(II) complex also corresponds to the ligand structure of L², in which one cyanomethyl group is attached to the sterically less hindered nitrogen atom. The electronic absorption spectra (Table 1) of [NiL²](ClO₄)₂·H₂O measured in Nujol mull and various solvents show a d-d transition band at *ca.* 477 nm. The wavelength is intermediate between those of the square-planar complexes [NiL¹](ClO₄)₂ (465 nm) and [NiL⁵](ClO₄)₂ (492 nm).^{17,25} This corresponds to the general trend that stepwise alkylation to coordinated nitrogen atoms of a 14-membered tetraaza macrocyclic nickel(II) complex weakens the ligand field strength.^{15,17,26} The visible absorption spectra of [CuL²](ClO₄)₂·H₂O measured in Nujol mull (495 nm) and nitromethane (497 nm) are also corresponding to the square-planar Cu-N₄ chromophore.^{15,17,25} Above results strongly indicate that [NiL²](ClO₄)₂·H₂O has a square-planer coordination geometry. The wavelengths measured in acetonitrile (527 nm) and DMSO (532 nm) are *ca.* 30 nm longer than those measured in Nujol mull and nitromethane,

implicating the coordination of the solvent molecule in the solvents.

The nickel(II) and copper(II) complexes are soluble in acetonitrile, nitromethane, or DMSO, but are nearly insoluble in methanol or water at room temperature. The complexes are quite inert against hydrolysis in acidic aqueous solutions (pH ≤ 6). As described above, however, they are readily hydrolyzed to [Ni(HL³)(H₂O)]²⁺ or [Cu(HL³)]²⁺ in basic aqueous solutions.

Spectra and Properties of [Ni(HL³)(H₂O)](ClO₄)₂, [Cu(HL³)](ClO₄)₂, and [NiL³(H₂O)]ClO₄. Infrared spectra of [Ni(HL³)(H₂O)](ClO₄)₂ and [Cu(HL³)](ClO₄)₂ show several peaks corresponding to $\nu_{\text{N-H}}$ of the coordinated amino and the amide groups at 3390-3190 cm⁻¹. The spectra also show $\nu_{\text{C=O}}$ of the pendant arm at *ca.* 1670 cm⁻¹. In the spectrum of [NiL³(H₂O)]ClO₄, four peaks of $\nu_{\text{N-H}}$ are observed at 3310-3200 cm⁻¹. The FAB mass spectra of the nickel(II) and copper(II) complexes show two groups of peaks corresponding to the fragments [M(HL³) + ClO₄]⁺ and [M(HL³) - H]⁺ (*see* Experimental section). The magnetic moments (μ_{eff}) of [Ni(HL³)(H₂O)](ClO₄)₂ and [NiL³(H₂O)]ClO₄ in the solid states are 2.80 and 2.78 μ_B , respectively, at room temperature. This is consistent with a *d*⁸ electronic configuration of the complexes in octahedral coordination geometry. The electronic absorption spectrum (Table 1) of [Ni(HL³)(H₂O)](ClO₄)₂ measured in Nujol mull shows three bands at *ca.* 340 (³B_{1g} → ³E_g(P)), 510 (³B_{1g} → ³E_g), and 690 (³B_{1g} → ³B_{2g} + ³B_{1g} → ³A_{2g}) nm, indicating that the complex has octahedral coordination geometry in the solid state;^{10,14,17,27} water molecule as well as the pendant amide group is involved in coordination. The spectra measured in various solvents are similar to that measured in Nujol mull. The spectra of [NiL³(H₂O)]ClO₄ measured in various solvents are comparable with those of [Ni(HL³)(H₂O)](ClO₄)₂. However, the

wavelength and molar absorption coefficient of each band for $[\text{NiL}^3(\text{H}_2\text{O})]\text{ClO}_4$ are somewhat longer and larger, respectively, than those for $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})](\text{ClO}_4)_2$. The electronic absorption spectra of $[\text{Cu}(\text{HL}^3)](\text{ClO}_4)_2$ measured in Nujol mull and various solvents show a d-d transition band at ca. 540 nm ($\epsilon = 125\text{--}130 \text{ M}^{-1}\text{cm}^{-1}$), supporting the suggestion that the complex has square-pyramidal coordination geometry.^{15,20} The copper(II) complex, unlike the nickel(II) complexes, is reluctant to form octahedral structure. This may be closely related to Jahn-Teller distortion of octahedral copper(II) complexes. The values of the molar conductance (Table 1) measured in various solvents indicate that $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})](\text{ClO}_4)_2$ (or $[\text{Cu}(\text{HL}^3)](\text{ClO}_4)_2$) and $[\text{NiL}^3(\text{H}_2\text{O})]\text{ClO}_4$ are 1:2 and 1:1 electrolytes, respectively. The deprotonated complex $[\text{NiL}^3(\text{H}_2\text{O})]\text{ClO}_4$ as well as $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})](\text{ClO}_4)_2$ and $[\text{Cu}(\text{HL}^3)](\text{ClO}_4)_2$ is quite stable in the solid state and in pure water, acetonitrile, or DMSO. Both $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})](\text{ClO}_4)_2$ and $[\text{Cu}(\text{HL}^3)](\text{ClO}_4)_2$ are also stable in acidic aqueous solutions.

Spectra and Properties of $[\text{NiL}^4(\text{H}_2\text{O})](\text{ClO}_4)_2$ and $[\text{CuL}^4](\text{ClO}_4)_2$. The FAB mass spectrum of $[\text{NiL}^4(\text{H}_2\text{O})](\text{ClO}_4)_2$ shows two groups of peaks at m/z 564.2 $\{[\text{NiL}^4 + \text{ClO}_4]^+\}$ and 464.3 $\{[\text{NiL}^4 - \text{H}]^+\}$. In the spectrum of $[\text{CuL}^4](\text{ClO}_4)_2$, two groups of peaks corresponding to $[\text{CuL}^4 + \text{ClO}_4]^+$ and $[\text{CuL}^4 - \text{H}]^+$ fragments are observed at m/z 569.1 and 470.2, respectively. Infrared spectra of the complexes show several peaks of $\nu_{\text{N-H}}$ at 3300–3200 cm^{-1} . The spectra also show $\nu_{\text{C=N}}$ of the coordinated *N*- $\text{CH}_2\text{C}(\text{=NH})\text{OCH}_3$ group at ca. 1665 cm^{-1} . The electronic absorption spectra (Table 1) of $[\text{NiL}^4(\text{H}_2\text{O})](\text{ClO}_4)_2$ measured in various solvents are comparable with those of other octahedral complexes, such as $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})](\text{ClO}_4)_2$, indicating that the water molecule as well as the pendant imidate ester group is involved in coordination. The magnetic moment (2.76 μ_{B}) of $[\text{NiL}^4(\text{H}_2\text{O})](\text{ClO}_4)_2$ is also consistent with a d^8 electronic configuration in octahedral coordination geometry. The spectra of $[\text{CuL}^4](\text{ClO}_4)_2$ measured in various solvents also similar to those of $[\text{Cu}(\text{HL}^3)](\text{ClO}_4)_2$ and other square-pyramidal complexes,^{15,20} indicating that the pendant imidate ester group is coordinated to the metal center. The nickel(II) and copper(II) complexes are quite stable in the solid state and in pure water, acetonitrile, or DMSO.

Solution Behaviors. Both $[\text{NiL}^4(\text{H}_2\text{O})]^{2+}$ and $[\text{CuL}^4]^{2+}$ readily undergo hydrolysis to give $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})]^{2+}$ or $[\text{Cu}(\text{HL}^3)]^{2+}$ in basic aqueous solutions (see above). It has been reported that the *N*- $\text{CH}_2\text{C}(\text{=NH})\text{OCH}_3$ pendant arms attached to $[\text{CuL}^7]^{2+}$ are readily hydrolyzed to *N*- $\text{CH}_2\text{COOCH}_3$ groups at $\text{pH} \leq 7$, though those attached to $[\text{NiL}^7]^{2+}$ are inert against hydrolysis under similar conditions.²³ However, $[\text{CuL}^4]^{2+}$ as well as $[\text{NiL}^4(\text{H}_2\text{O})]^{2+}$ was found to be quite inert against hydrolysis in neutral or acidic aqueous solutions ($1 \leq \text{pH} \leq 7$); no apparent hydrolysis was observed even after 24 h at room temperature. The stability of $[\text{CuL}^4]^{2+}$ and $[\text{NiL}^4(\text{H}_2\text{O})]^{2+}$ in acidic aqueous solutions was also confirmed by the recrystallization of the complexes from warm 0.1 M HClO_4 aqueous solutions.

Electronic absorption spectrum of $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})]^{2+}$ or

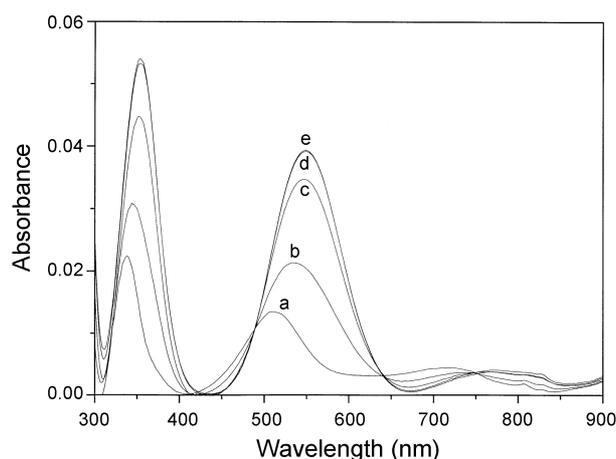


Figure 1. Visible absorption spectra of $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})](\text{ClO}_4)_2$ ($2.0 \times 10^{-3} \text{ M}$) measured in NaOH water-acetonitrile (1:1) solutions at 25 °C; pH 6.6 (a), 10.4 (b), 11.1 (c), 12.2 (d), and 12.7 (e). The spectra measured in 0.1 M HClO_4 and 0.1 M NaOH solutions are similar to curves a and e, respectively.

$[\text{Cu}(\text{HL}^3)]^{2+}$ measured in 0.1 M HClO_4 solution was found to be nearly the same as that measured in neutral aqueous solution, indicating that the *N*- CH_2CONH_2 pendant arm is not protonated in the solution. Figure 1 shows that the addition of NaOH to an aqueous solution of $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})]^{2+}$ shifts the absorption bands at 337 and 506 nm to longer wavelengths and increases their molar absorption coefficients. This is attributed to the fact that the complex exists as an equilibrium mixture of $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})]^{2+}$ and $[\text{NiL}^3(\text{H}_2\text{O})]^+$ (Eq. (1)) in basic aqueous solutions. Figure 1 also shows that the majority part of the complex exists as $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})]^{2+}$ at $\text{pH} \leq 10$, whereas most of the complex is deprotonated to form $[\text{NiL}^3(\text{H}_2\text{O})]^+$ at $\text{pH} \geq 11$. Although the formation of $[\text{NiL}^3(\text{H}_2\text{O})]^+$ was confirmed by the isolation of $[\text{NiL}^3(\text{H}_2\text{O})]\text{ClO}_4$, the deprotonation of the coordinated water could not be excluded in the basic aqueous solutions. The spectra (Fig. 2) of $[\text{Cu}(\text{HL}^3)]^{2+}$ measured in NaOH aqueous solutions also

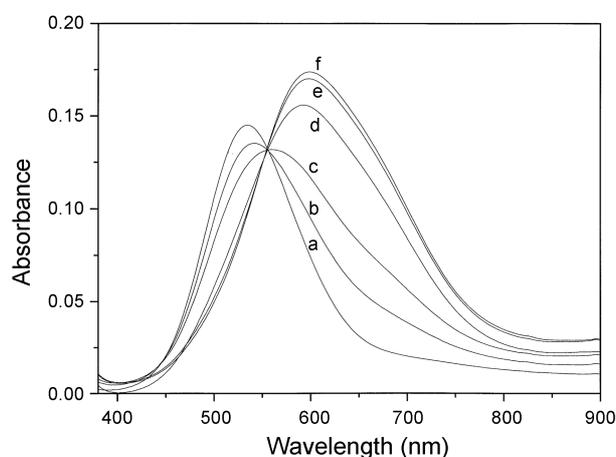
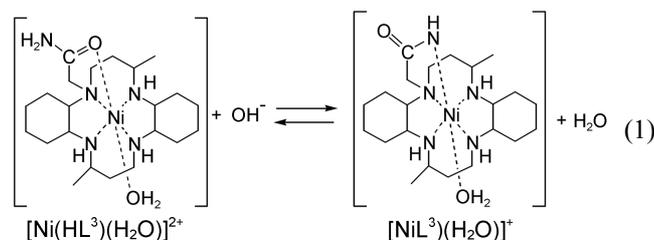


Figure 2. Visible absorption spectra of $[\text{Cu}(\text{HL}^3)](\text{ClO}_4)_2$ ($1.0 \times 10^{-3} \text{ M}$) measured in NaOH water-acetonitrile (1:1) solutions at 25 °C; pH 7.0 (a), 11.4 (b), 12.0 (c), 12.7 (d), $[\text{NaOH}] = 0.05 \text{ M}$ (e), and $[\text{NaOH}] = 0.1 \text{ M}$ (f). The spectrum measured in 0.1 M HClO_4 solution is similar to curve a.

show that the addition of NaOH shifts the band to a longer wavelength. This also supports the deprotonation of the copper(II) complex and/or the coordination of water (or hydroxide ion) in basic aqueous solutions. In contrast to the case of $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})]^{2+}$, however, the only copper(II) complex isolated as a solid in the basic aqueous solutions was $[\text{Cu}(\text{HL}^3)](\text{ClO}_4)_2$.

The approximate pK_a values of $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})]^{2+}$ and $[\text{Cu}(\text{HL}^3)]^{2+}$ were determined at 25 °C by using a spectrophotometric method.²⁸ The pK_a value (*ca.* 10.6) for $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})]^{2+}$ was found to be much smaller than that (≥ 12.2) for $[\text{Cu}(\text{HL}^3)]^{2+}$. One of the reasons for the easier deprotonation of $[\text{Ni}(\text{HL}^3)(\text{H}_2\text{O})]^{2+}$, compared to that of $[\text{Cu}(\text{HL}^3)]^{2+}$, may be the stronger interaction between the *N*-CH₂CONH⁻ group and the central metal ion of the conjugate base ($[\text{NiL}^3(\text{H}_2\text{O})]^+$).¹⁵ In the case of the copper(II) complex, the interaction between the functional group and the metal ion is expected to be relatively weak because of the Jahn-Teller distortion.



Concluding Remarks

This work shows that the mono-*N*-cyanomethylated macropolycycle L¹¹, which can be prepared by the reaction of bromoacetonitrile with L¹⁰ containing one *N*-CH₂-*N* linkage, is a useful precursor for the preparation of various types of mono-*N*-functionalized macrocyclic compounds, such as the nickel(II) and copper(II) complexes of L²-L⁴. This work also shows that chemical properties of the *N*-CH₂CONH₂ or *N*-CH₂C(=NH)OCH₃ group attached to a macrocyclic complex is strongly influenced by the number of the functional pendant arm and by the nature of the central metal ion.

Acknowledgments. This work was supported by the grant of Daegu University (2010).

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