# A Novel Synthesis and Crystal Structure of 2,3-Substituted-1,4-2H-tetrazolthione 

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The tetrazole derivatives have been extensively studied due to their explosive properties and biological importances in agriculture, biochemistry and pharmacology. ${ }^{1}$ Tetrazolthiones show antibacterial activity and they are known as good weed killers. ${ }^{2}$ In general, tetrazolthione derivatives were synthesized by the reactions of alkyl isothiocyanates with azides, or by the reactions of alkyl isoaminothiourea with nitrous acid. However, only 1,4 -substituted tetrazolthiones could be prepared in these methods. ${ }^{3}$ In our laboratory, we found a new ring-closing reaction of dithizone with carbon disulfide to afford 2,3-disubstituted tetrazolthiones in high yields. By this method, 2,3-diphenyl-1,4-2H-tetrazolthione (1) and 2,3-di(p-methyl-phenyl)-1,4-2H-tetrazolthione (2) were synthesized and characterized by X-ray crystallography. In this paper, we describe the synthesis of the two compounds and the crystal structure of 2.

## Experimental Section

All chemicals were obtained from a commercial source and used without additional purification.
Synthesis. Dithizone ( $2.0 \mathrm{~g}, 8.0 \mathrm{mmol}$ ) was dissolved in acetonitrile ( 80 mL ). To this solution was added the mixture of carbon disulfide ( 16.0 mmol ) and $50 \%$ aqueous sodium hydroxide ( 8.0 mmol ) with stirring at $40^{\circ} \mathrm{C}$. The reaction mixture was kept at $40^{\circ} \mathrm{C}$ for 4 h to form red precipitates, and then it is cooled to room temperature. The red crystalline solids were collected by filtration and recrystallize from EtOH to give 1. Yield: $86 \%$, mp $167-168{ }^{\circ} \mathrm{C}$. Calc. for $\mathrm{C}_{13} \mathrm{H}_{12} \mathrm{~N}_{4} \mathrm{~S}: \mathrm{C}, 60.91$; H, 4.71 ; N, 21.86\%. Found: C, 60.85; $\mathrm{H}, 4.67$; N, 21.79\%. IR (KBr), $v_{\max }: 3440(\mathrm{~N}-\mathrm{H}), 3010\left(\mathrm{C}_{\mathrm{Ar}}-\right.$
H), 1588, 1487, $1463\left(\mathrm{C}_{\mathrm{Ar}}-\mathrm{C}_{\mathrm{Ar}}\right), 1315(\mathrm{C}=\mathrm{S}), 1244$ ( $\mathrm{N}-\mathrm{N}$,


Compound 2 was prepared with the same procedure as described for $\mathbf{1}$ except that $p, p^{\prime}$-dimethyl-dithizone ( 2.3 g , 8.0 mmol ) was used to replaced the dithizone. Yield $90 \%$. $\mathrm{mp} 169-171{ }^{\circ} \mathrm{C}$. Calc. for $\mathrm{C}_{15} \mathrm{H}_{16} \mathrm{~N}_{4} \mathrm{~S}$ : C, 63.29; H, 5.67 ; N, $19.71 \%$. Found: C, 63.20; H, 5.70; N, 19.64\%. IR (KBr), $\nu_{\text {max }}: 3432(\mathrm{~N}-\mathrm{H}), 3027\left(\mathrm{C}_{\mathrm{Ar}}-\mathrm{H}\right), 2990,2920\left(\mathrm{C}_{\text {methyl }}-\mathrm{H}\right)$, 1506, $1401\left(\mathrm{C}_{\mathrm{Ar}}-\mathrm{C}_{\mathrm{Ar}}\right), 1296(\mathrm{C}=\mathrm{S}), 1241\left(\mathrm{~N}-\mathrm{N}, \mathrm{C}_{\mathrm{Ar}}-\mathrm{N}\right),{ }^{5,6}$ 980, 826, $712\left(\mathrm{C}_{\mathrm{Ar}^{-}}-\mathrm{H}\right) \mathrm{cm}^{-1}$.

The synthetic pathway is shown in Scheme 1.
X-ray Structures of 2. The selected crystals of $\mathbf{2}$ were mounted on a glass fiber. The data were collected with graphic monochromated $\mathrm{Mo}-\mathrm{K} \alpha(\lambda=0.71073 \AA)$ radiation at 293 K . The collected data were reduced by using the program SAINT and the empirical absorption correction was done by using the $S A D A B S$ program. The structure was solved by direct method and refined by full-matrix leastsquares method on $F_{\text {obs }}{ }^{2}$ by using the SHELXTL program. All non-H atoms were anisotropically refined. The hydrogen atoms were located by difference synthesis and refined geometrically. Final conventional $R_{1}=0.1089, w R_{2}=0.3064$, $\mathrm{S}=1.219$.

## Results and Discussion

Compound 2 was crystallized in the orthorhombic system, space group $\mathrm{Pmn} 2_{l}$ with $a=18.381$ (4), $b=6.180$ (1), 6.18 7(1) $\AA, \mathrm{C}_{15} \mathrm{H}_{16} \mathrm{~N}_{4} \mathrm{~S}, M_{r}=284.38, V=702.7$ (2) $\AA^{3}, Z=2, D_{c}=$ $1.344 \mathrm{~g} / \mathrm{cm}^{3}, F(000)=300, \mu=0.226 \mathrm{~mm}^{-1}$. The OPTEP drawing with the numbering scheme for 2 is shown in Figure 1.


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Figure 1. Molecular structure with the atomic numbering scheme for 2.

The crystal structure consists of two 2,3-di(p-methyl-phenyl)-1,4-2H-tetrazolthione molecules. The bond lengths and angles in the phenyl ring are generally normal. The $\mathrm{N}(1)$ atom, $\mathrm{C}(7)$ atom and phenyl ring fall within a plane. The plane equation is $14.6972 x+2.6120 y-2.6395 z=0.7440$, with the largest atom deviation is $0.019 \AA$ The bond lengths and angles in the tetrazole ring are comparable to those reported before. ${ }^{6,7}$ The $\mathrm{N}(1)-\mathrm{N}(2)$ bond distance, 1.313(8) $\AA$, is indicative of some double-bond character. The leastsquares planes of phenyl ring and tetrazole ring are almost perpendicular, with the dihedral angle being $89.65^{\circ}$. The bond length of $\mathrm{S}(1)-\mathrm{C}(8), 1.686$ (10) $\AA$, is in the normal range.

In the crystal lattice, there exists a $\mathrm{C}-\mathrm{H} \cdots \pi$ supramolecule interaction. ${ }^{8,9}$ The distance between $\mathrm{C}(7)-\mathrm{H}(7 \mathrm{~A})$ to phenyl ring is $3.030 \AA$, which stabilizes the crystal structure.

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Supporting Information Materials. Listing of atomic coordinates, complete bond distances and angles, thermal parameters, and least-squares results for the title compounds $\mathbf{1}$ and $\mathbf{2}$ are available on request from the corresponding author.

## References

1. Katritzky, A. R.; Ress, C. W. Comprehensive Heterocyclic Chemistry, 1st Ed; Pergamon Press: England, 1984; Vol 5, Part 4A.
2. Assy, M. G. Pol. J. Chem. 1995, 69, 1022.
3. Lieber, E.; Ramachandran, J. Can. J. Chem. 1959, 37, 101.
4. Wu, J. G. Technology and Application of Fourier Transformation Infrared Spectroscopy; Science and Technology Publishing Company: Beijing, 1994.
5. Jha, J. K.; Kumar, A.; Ojha, V. S. J. Indian Chem. Soc. 1997, 72, 676.
6. Hanley, R. N.; Ollis, W. D.; Ramsden, C. A.; Smith, I. S. J. Chem. Soc., Perkin II 1979, 744.
7. Preston, P. N.; Tiwan, K. K.; Turnbull, K. J. Chem. Soc., Chem. Commun. 1976, 343.
8. Glusker, J. P.; Lewis, M.; Rossi, M. Crystal Structure Analysis for Chemists and Biologists; VCH Publishers Inc.: New York: 1994.
9. Hunter, R. H.; Haueisen, R. H.; Irving, A. Angew. Chem. Int. Ed. Engl. 1994, 33, 566.

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