

## Microwave-assisted Kabachnik-Fields Reaction in Ionic Liquid

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In recent years one of the prime concerns of chemists is search for replacements of the conventional organic solvents with environmentally benign ones. In this context, room temperature ionic liquids are increasing interest as reaction media. In particular, utilization of the ionic liquids in metal-catalyzed reactions allowed easy recovery and reuse of the catalysts.<sup>1</sup> During our on going study on catalyst recycling using ionic liquids,<sup>2</sup> we interested in the combination of ionic liquids with microwave (MW)-assisted organic reactions.<sup>3</sup> Ionic liquids consisting entirely of ions could absorb the microwave energy efficiently. Moreover, the risk of an explosion caused by rapid increase of vapor pressure of the solvent can be avoided since ionic liquids have no measurable vapor pressure. In spite of several advantages, only few MW-assisted reactions have been carried out in ionic liquids media and no successful catalyst recycling has been achieved so far.<sup>4</sup> Here we wish to report our preliminary results on the application of ionic liquids in a MW-assisted reaction.

To demonstrate the utility of ionic liquids for catalyst recycling in MW-assisted metal-catalyzed organic reactions, lanthanide triflates catalyzed Kabachnik-Fields reaction<sup>5</sup> was chosen as a model reaction. Kabachnik-Fields reaction is one of the effective methods for the synthesis of biologically important  $\alpha$ -aminophosphonates and has been received a great deal of attention in recent years.<sup>6</sup> Thus, a solution of benzaldehyde, aniline and diethylphosphite in [bmim][PF<sub>6</sub>] in the presence of 10 mol% of Yb(OTf)<sub>3</sub> as a catalyst was irradiated for 20 s using a domestic MW oven (2450 MHz, 850W), then the reaction mixture was taken out, mixed again for 5 s and then irradiated again for 20 s. This step was repeated until completion of the reaction (*ca.* 2 min) to give product in 86% yield. Surprisingly, when the solvent was changed to [bmim][BF<sub>4</sub>], the yield was markedly increased, thus, the product was obtained almost quantitatively. However, the yield decreased under thermal condition (120 °C, 2 h, 74%) or under MW condition in other ionic liquids (55% in [bmim][OTf], 72% in [bmim][SbF<sub>6</sub>]) and DMF (70%) (entry 1 in Table 1). The catalytic activities of other catalysts were also examined under MW irradiation (Table 1).

As shown in Table 1, all of the catalysts examined exhibited superior catalytic activities in [bmim][BF<sub>4</sub>] and [bmim][PF<sub>6</sub>]. For examples, Yb(OTf)<sub>3</sub> and Dy(OTf)<sub>3</sub> showed higher activity in [bmim][BF<sub>4</sub>] than in other ionic liquids. On the other hand, the catalytic activities of the other catalysts, Sc(OTf)<sub>3</sub>, Gd(OTf)<sub>3</sub> and InCl<sub>3</sub>, are superior in [bmim]

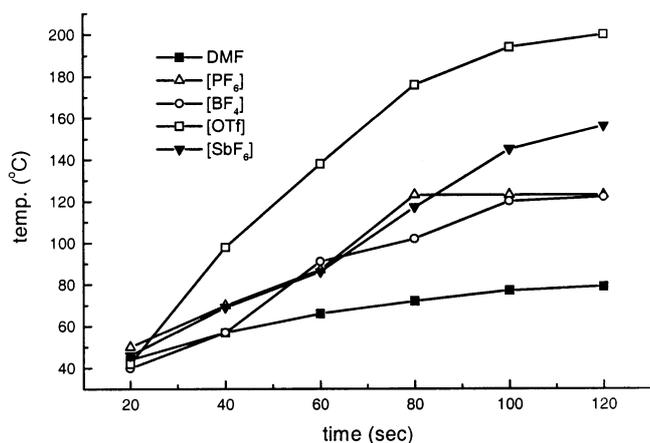
**Table 1.** The three component reactions of benzaldehyde, aniline and diethylphosphite in [bmim][X] or DMF in the presence of lanthanide triflates and indium trichloride using microwave<sup>a-c</sup>

Entry	catalyst	X in [bmim][X]				DMF
		PF <sub>6</sub>	BF <sub>4</sub>	OTf	SbF <sub>6</sub>	
1	Yb(OTf) <sub>3</sub>	86	99(74) <sup>d</sup>	55	72	70
2	Sc(OTf) <sub>3</sub>	88	81	80	63	90
3	Dy(OTf) <sub>3</sub>	87	90	55	80	62
4	Gd(OTf) <sub>3</sub>	94	62	52	63	65
5	InCl <sub>3</sub>	91	80	70	65	82

<sup>a</sup>Aldehyde (1 mmol), aniline (1.2 mmol), and diethyl phosphite (1.2 mmol) were successively added to a solution of catalyst (10 mole %) in an ionic liquid (1 mL) placed in a flat bottom flask, mixed thoroughly on a vortex Mixer (Scientific Inc., Model G-560) and the mixture was irradiated intermittently in an unmodified domestic microwave oven. After completion of the reaction, the product was isolated by extraction with benzene (5 mL  $\times$  4), and purified by silica column chromatography. For catalyst recycling, more reactant was added to the ionic liquid containing catalyst. <sup>b</sup>MW irradiated for 2 min (20 s + 20 s + 20 s + 20 s + 20 s). <sup>c</sup>Isolated yield (average of two runs). <sup>d</sup>Yield given in parenthesis obtained under thermal condition (120 °C, 2 h)

[PF<sub>6</sub>]. Moreover, higher or comparable catalytic activities were observed in ionic liquids compared to the organic solvent, DMF. In the recycling study, the catalyst, Yb(OTf)<sub>3</sub>, immobilized in [bmim][BF<sub>4</sub>] could be re-used 5 times without any loss of catalytic activity (>99% average yield). This result clearly indicates that both [bmim][BF<sub>4</sub>] solvent and Yb(OTf)<sub>3</sub> catalyst are quite stable upon microwave irradiation. On contrary, the catalytic activity of the same catalyst immobilized in [bmim][SbF<sub>6</sub>] was dramatically decreased in the second cycle (33% yield). The different catalytic properties of Yb(OTf)<sub>3</sub> in [bmim][BF<sub>4</sub>] and [bmim][SbF<sub>6</sub>] may be attributed to the different thermal behavior of the ionic liquids upon MW irradiation. Figure 1 shows temperatures attained as a function of irradiation time for ionic liquids, [bmim][X], and DMF.

It has been found that the heating rates of the ionic liquids, [bmim][X], are faster than DMF, and largely affected by the counter anion, X. The [bmim][OTf] showed the fastest heating rate, thus, the temperature reached 200 °C after 2 min. It is interesting that the temperature of [bmim][PF<sub>6</sub>] and



**Figure 1.** Thermal behaviors of ionic liquids and DMF under MW irradiation. Temperature is the mean average ( $\pm 5$  °C error range) of three measurements.

**Table 2.** Microwave assisted three component reactions of aldehydes, aniline and diethylphosphite in [bmim]BF<sub>4</sub> in the presence of Yb(OTf)<sub>3</sub> as catalyst

$$\text{Ar-CHO} + \text{Ph-NH}_2 + \text{H(O)P(OEt)}_2 \xrightarrow[\text{MW (2 min)}]{\text{[bmim][BF}_4\text{]}} \text{Ar-CH(NHPh)P(OEt)}_2$$

(10 mol %) Yb(OTf)<sub>3</sub>

Entry	Ar	Yield (%) <sup>a</sup>
1	Ph	>99
2	<i>p</i> -MeO-Ph	>99
3	<i>o</i> -Me-Ph	>99
4	<i>m</i> -Me-Ph	>99
5	<i>p</i> -Me-Ph	>99
6	<i>p</i> -F-Ph	94
7	1-Np	83

<sup>a</sup>Isolated yield

[bmim][BF<sub>4</sub>] reached thermal stationary states after 100 sec (*ca.* 115–120 °C). Although it is not clear yet, it may be possible to assume that the high activity and stability of Yb(OTf)<sub>3</sub> in [bmim][BF<sub>4</sub>] is largely related with its thermal behavior under MW irradiation. Finally, numbers of  $\alpha$ -

aminophosphonates could be rapidly synthesized in excellent yields through MW assisted Kabachnik-Fields reactions and the results are summarized in Table 2.

In summary, using lanthanide triflates catalyzed Kabachnik-Fields reaction as a model reaction has proved the utility of ionic liquids, [bmim][X], as solvents in MW-assisted metal-catalyzed reactions. Under MW irradiation, the thermal behaviors of ionic liquids and the recyclability of the catalysts were largely affected by the counter anions of [bmim][X]. A successful recycling of a catalyst, Yb(OTf)<sub>3</sub>, immobilized in [bmim][BF<sub>4</sub>] may open new eco-friendly protocol in MW-assisted metal-catalyzed reactions. Further applications of ionic liquids to other MW-assisted metal-catalyzed reactions are under study.

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## References

- (a) Welton, T. *Chem. Rev.* **1999**, *99*, 2071. (b) Wasserscheid, P.; Keim, W. *Angew. Chem., Int. Ed.* **2000**, *39*, 3772. (c) Brown, R. A.; Pollet, P.; McKoon, E.; Eckert, C. A.; Liotta, C. L.; Jessop, P. G. *J. Am. Chem. Soc.* **2001**, *123*, 1254.
- (a) Lee, S.-g.; Park, J. H.; Kang, J.; Lee, J. K. *Chem. Commun.* **2001**, 1698. (b) Song, C. E.; Shim, W. H.; Roh, E. J.; Lee, S.-g.; Choi, J. H. *Chem. Commun.* **2001**, 1122. (c) Song, C. E.; Oh, C. R.; Roh, E. J.; Choo, D. J. *Chem. Commun.* **2000**, 1743.
- Recent selected reviews for MW-assisted reactions. See (a) Perreux, L.; Loupy, A. *Tetrahedron* **2001**, *57*, 9199. (b) Lidström, P.; Tierney, J.; Wathey, B.; Westman, J. *Tetrahedron* **2001**, *57*, 9225.
- (a) Dubreuil, J. F.; Bazureau, J. P. *Tetrahedron Lett.* **2000**, *41*, 7351. (b) Dubreuil, J. F.; Bazureau, J. P. *Tetrahedron Lett.* **2001**, *42*, 6097. (c) Varma, R. S.; Nambodiri, V. V. *Pure Appl. Chem.* **2001**, *73*, 1309.
- (a) Kabachnik, M. J.; Medved, T. *Izv. Akad. Nauk SSSR* **1954**, 1126. (b) Filds, E. K. *J. Am. Chem. Soc.* **1952**, *74*, 1528.
- (a) Recent papers for the synthesis of  $\alpha$ -aminophosphonates. See: Qian, C.; Huang, T. *J. Org. Chem.* **1998**, *63*, 4125. (b) Ranu, B. C.; Hajira, A.; Jana, U. *Org. Lett.* **1999**, *1*, 1141. (c) Manabe, K.; Kobayashi, S. *Chem. Commun.* **2000**, 669. (d) Kaboudin, B.; Nazari, R. *Tetrahedron Lett.* **2001**, *42*, 8211. (e) Chandrasekhar, S.; Prakash, S. J.; Jagadeshwar, V.; Narsihmulu, C. *Tetrahedron Lett.* **2001**, *42*, 5561.